



Lakshya JEE (2024)

Salt Analysis

DPP-03

- $\text{Na}_2\text{S} + \text{HCl} \longrightarrow \text{NaCl} + \text{"X"}$. X is:
 (1) H_2S (2) H_2SO_4
 (3) H_2SO_3 (4) $\text{H}_2\text{S}_2\text{O}_3$
- Which of the following gas has rotten egg smell
 (1) PH_3 (2) NH_3
 (3) H_2S (4) PbS
- $\text{H}_2\text{S} + \text{Na}_2[\text{Fe}(\text{CN})_5\text{NO}] \longrightarrow$ No reaction. It means
 (1) S^{2-} ion absent
 (2) S^{2-} ion may be present
 (3) S^{2-} ion is present in excess
 (4) S^{2-} ion form precipitate
- $\text{X} + \text{ZnCO}_3/\text{CdCO}_3 \longrightarrow$ ppt. It shows that 'X' is:
 (1) Cl^- ion (2) N^{3-} ion
 (3) S^{2-} ion (4) $\text{S}_2\text{O}_3^{2-}$ ion
- $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O} \xrightarrow{\Delta} \text{"X"}$, where X is:
 (1) Na_2S_5 (2) Na_2SO_4
 (3) $\text{Na}_2\text{S}_2\text{O}_3$ (4) No reaction.
- $\text{"X"} + \text{AgNO}_3 \longrightarrow [\text{Ag}(\text{S}_2\text{O}_3)_2]^{3-}$;
 $\text{"X"} + \text{AgNO}_3 \xrightarrow[\text{(excess)}]{\Delta} \text{Ag}_2\text{S}$. Then, X can be
 (1) $\text{S}_2\text{O}_3^{2-}$ (2) S^{2-}
 (3) SO_4^{2-} (4) Cl^-
- Acetate gives deep red colour on reaction with—
 (1) Ferrous Chloride (2) Ferric Chloride
 (3) Ferrous acetate (4) Ferric acetate
- The colour of the solution in product of the given reaction is
 $\text{CH}_3\text{COONa} + \text{FeCl}_3 + \text{H}_2\text{O} \rightarrow$
 (Neutral)
 (1) Reddish Brown (2) Yellow
 (3) Black (4) Green
- A substance on treatment with dil. H_2SO_4 liberates a colourless gas which produces
 (i) Turbidity with baryta water and
 (ii) Turns acidified dichromate solution green
 The reaction indicates the presence of
 (1) $\text{C}_2\text{O}_3^{2-}$ (2) S^{2-}
 (3) SO_3^{2-} (4) NO_2^-
- What is possible oxidation state of S in $\text{Na}_2\text{S}_2\text{O}_3$?
 (1) 0 (2) +4
 (3) -2 (4) +6



Note: Kindly find the Video Solution of DPPs Questions in the DPPs Section.

Answer Key

1. (1)
2. (3)
3. (2)
4. (3)
5. (3)

6. (1)
7. (2)
8. (1)
9. (3)
10. (3, 4)



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